

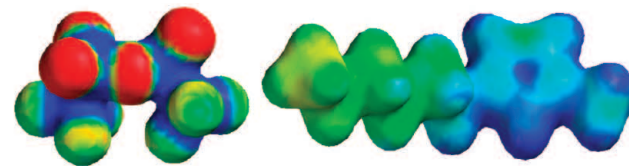
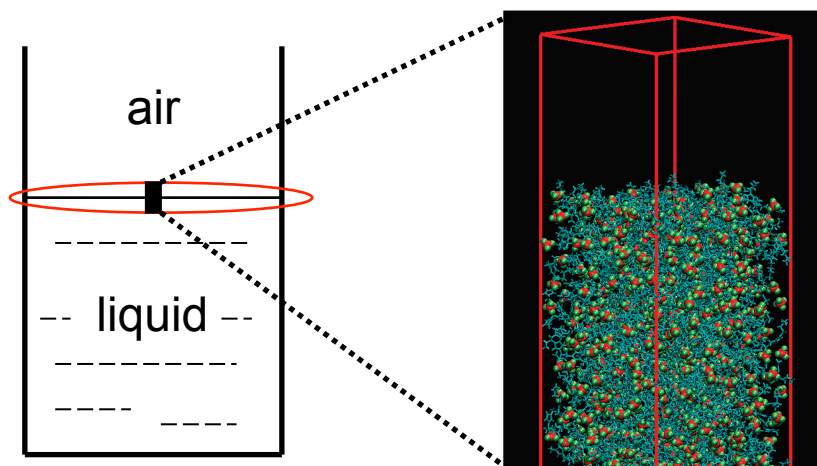


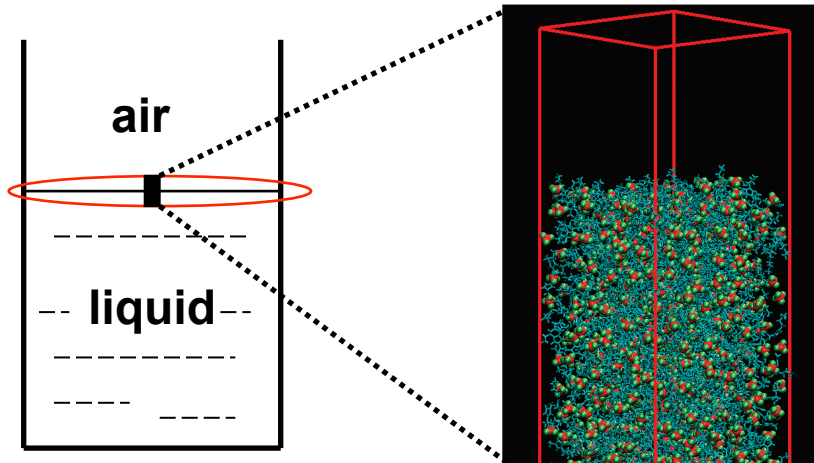
# Vacuum-Liquid Interface of Imidazolium-Based Ionic Liquids from Atomistic Simulations and High-Resolution Rutherford Backscattering Spectroscopy

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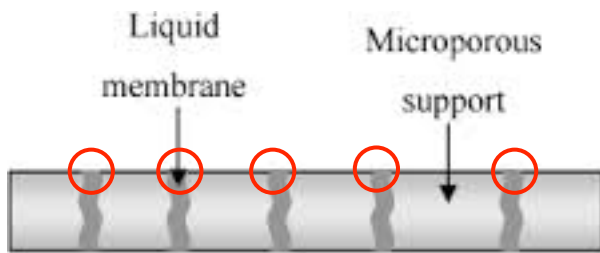
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- Fluid-fluid interfaces play a key role in various physical and chemical processes (adsorption, liquid-liquid extraction, heterogeneous catalysis, chromatography, or drug delivery).
- Physical properties such as density, solubility or dielectric constant change abruptly, yet continuously, in the vicinity of the interfaces.
- The molecules located at the interface experience a markedly different local environment wrt bulk phases.
- Single-particle properties such as orientation, diffusion constant or binding energy of the interfacial molecules are markedly different than those of the bulk phases.

## *Molecular Transport across Liquid Membrane*

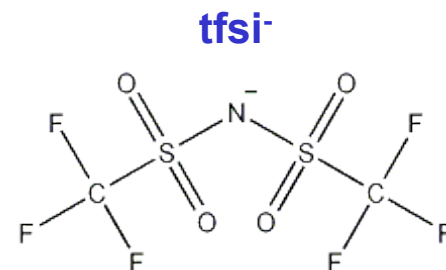
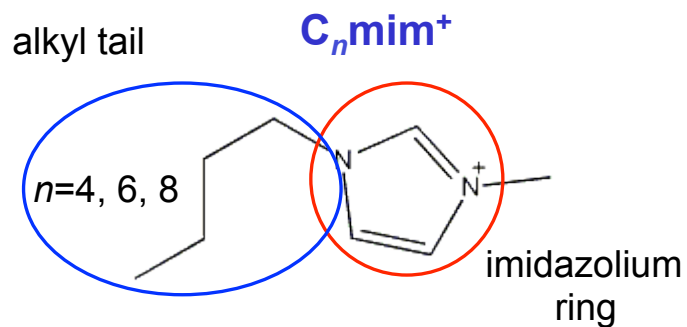


- solubility of molecules
- diffusion
  - diffusion in bulk phase
  - **interfacial transport – mass transfer resistance**



- Size of the interfacial region?
  - Ions density and concentration at the interfacial region?
  - Any preferential orientation of ions at the interfacial region?
  - Relation of the interfacial ions structure with surface tension?
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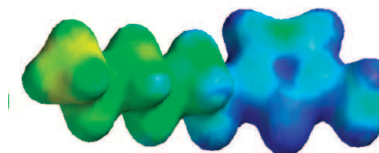
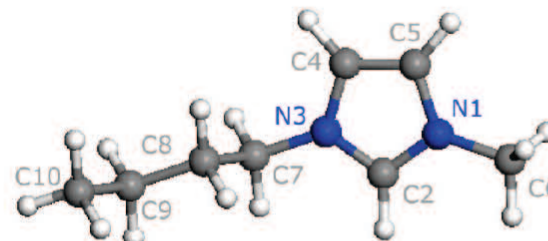
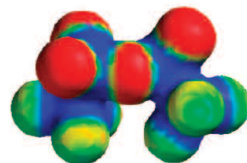
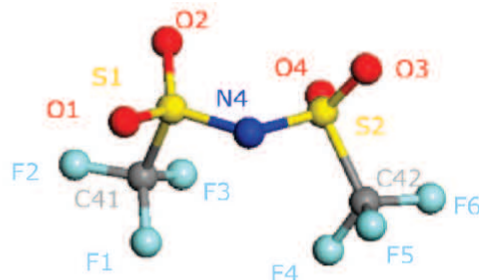
## Imidazolium-Based Ionic Liquids





## Flexible All-Atom Models

[C<sub>4</sub>mim][tfsi]



## Thermodynamic Properties @ T=300K & P=1bar

| RTILs                      | $\rho^{sim}$<br>(kg/m <sup>3</sup> ) | $\rho^{exp}$<br>(kg/m <sup>3</sup> ) | $\Delta\rho$<br>(%) | $\Delta H_v^{sim}$<br>(kJ/mol) | $\Delta H_v^{exp}$<br>(kJ/mol) | $\Delta H_v$<br>(%) |
|----------------------------|--------------------------------------|--------------------------------------|---------------------|--------------------------------|--------------------------------|---------------------|
| [C <sub>4</sub> mim][tfsi] | 1461.6(55)                           | 1436.2(6)                            | 1.7                 | 133.1(20)                      | 134(3)                         | -0.1                |
| [C <sub>6</sub> mim][tfsi] | 1394.9(48)                           | 1364.0(26)                           | 2.2                 | 140.2(32)                      | 139(2)                         | 0.9                 |
| [C <sub>8</sub> mim][tfsi] | 1338.5(42)                           | 1320.0(50)                           | 1.4                 | 147.0(38)                      | 149(2)                         | -1.3                |

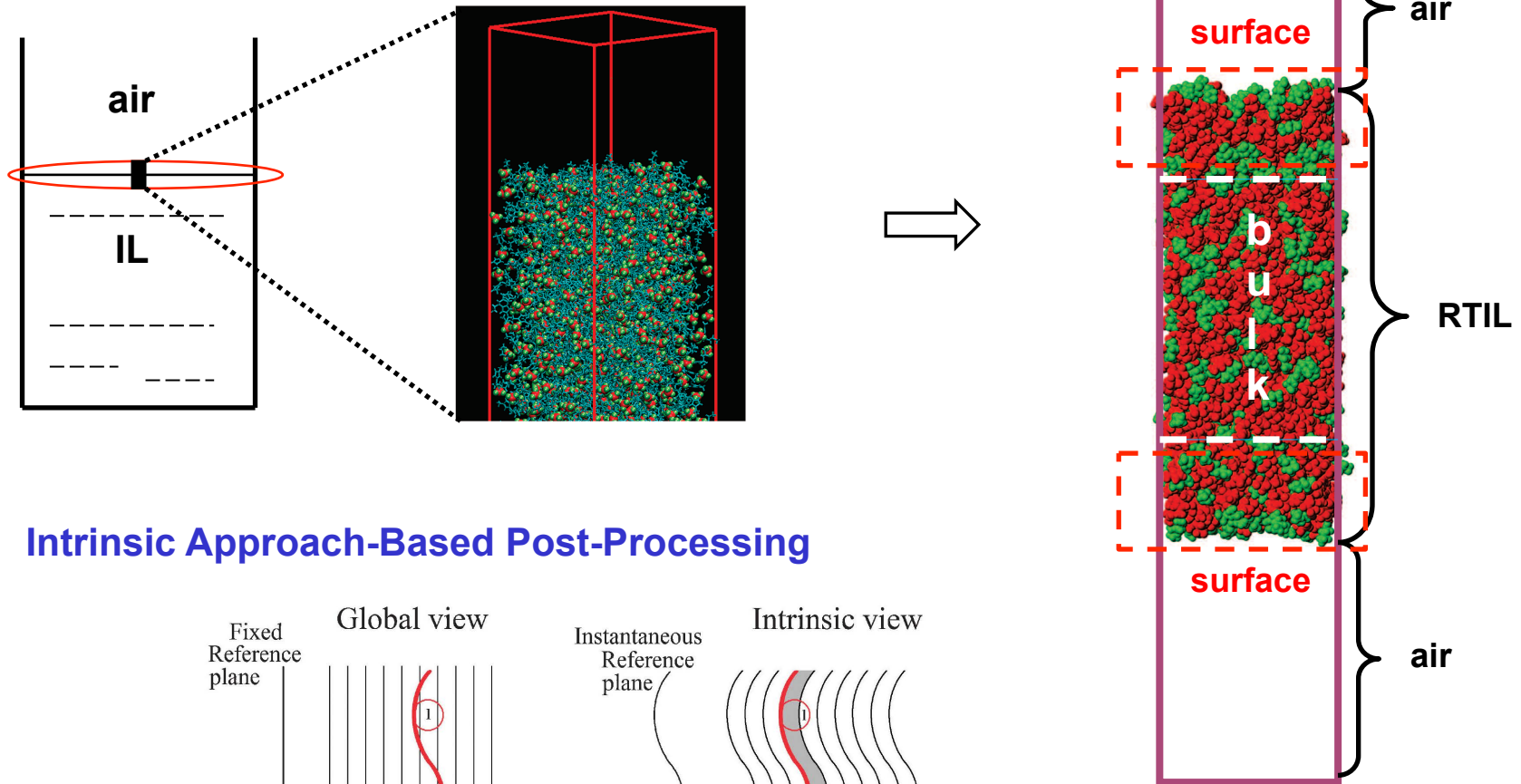




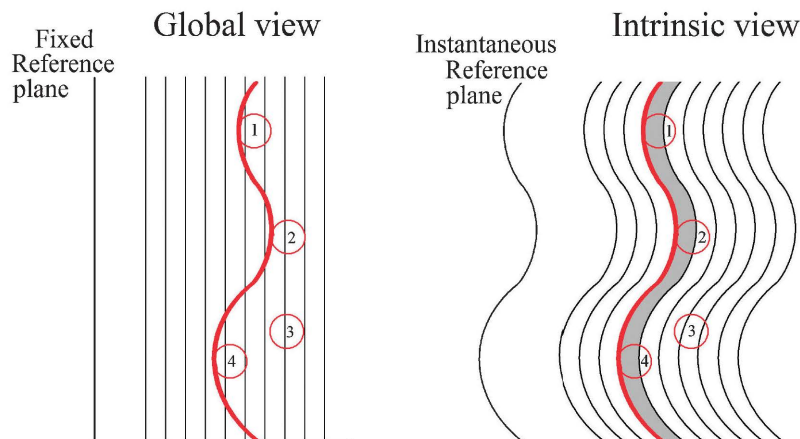
# Molecular Dynamics Simulations



## Simulation Set-up



## Intrinsic Approach-Based Post-Processing

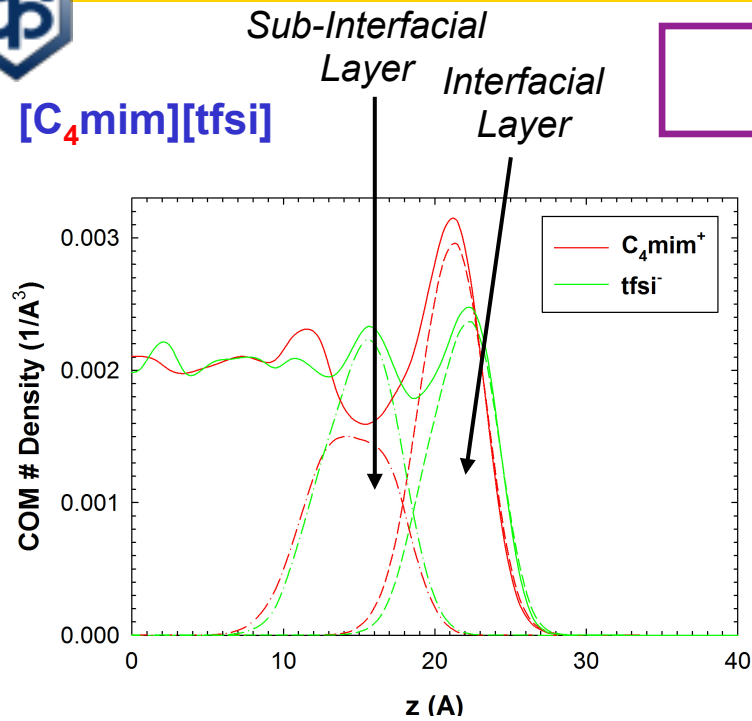




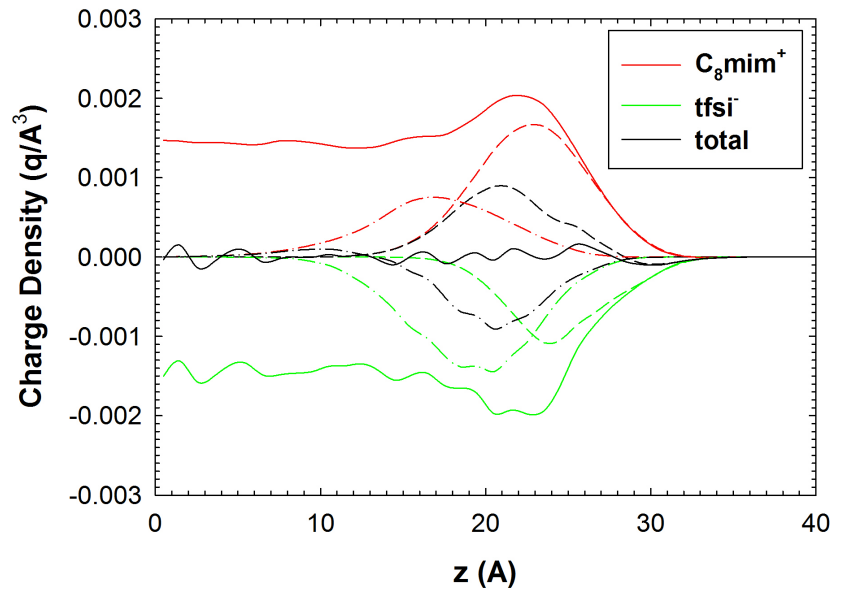
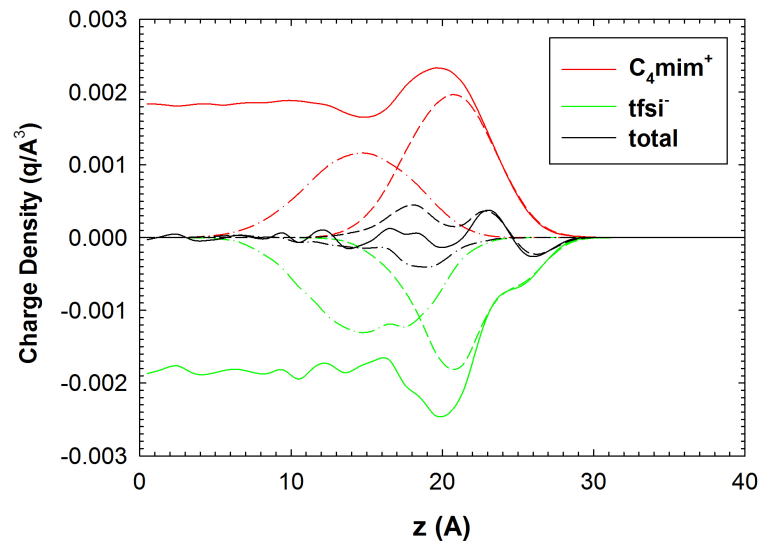
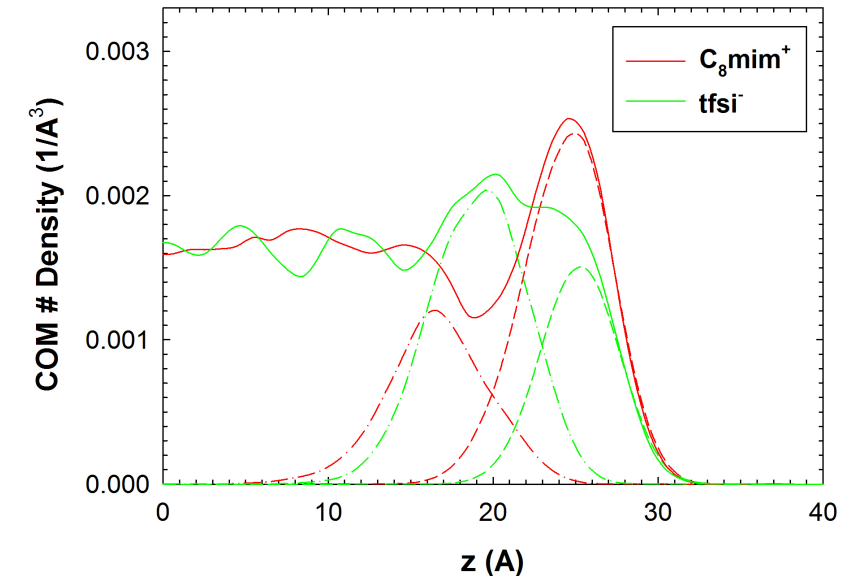
# Ions Density Profiles



**[C<sub>4</sub>mim][tfsi]**



**[C<sub>8</sub>mim][tfsi]**

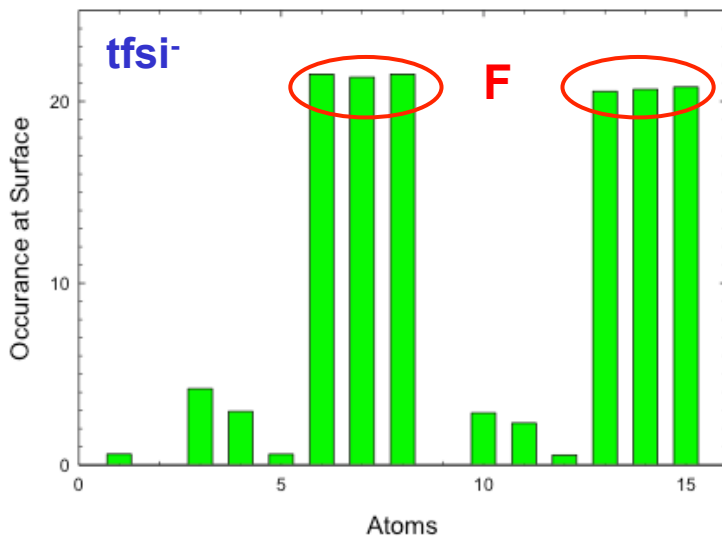
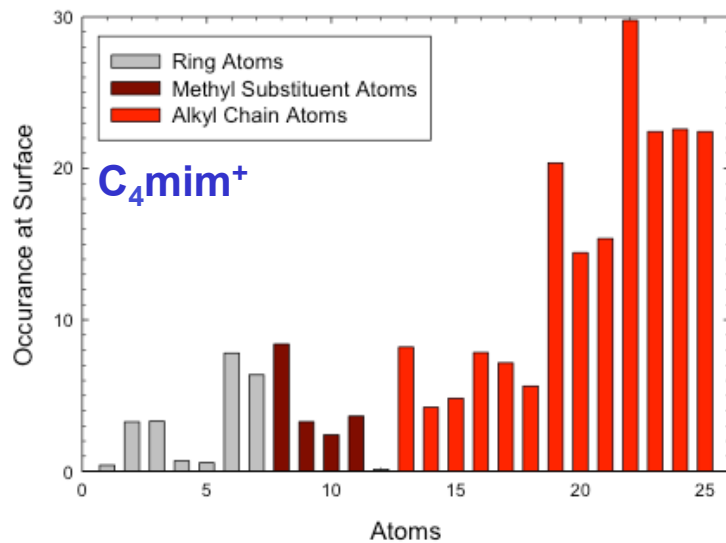




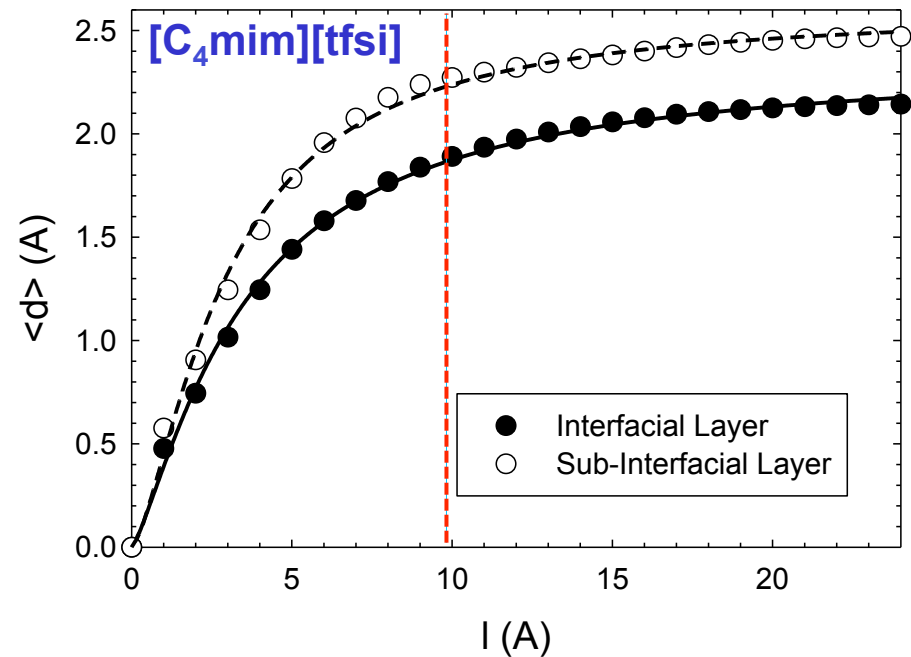
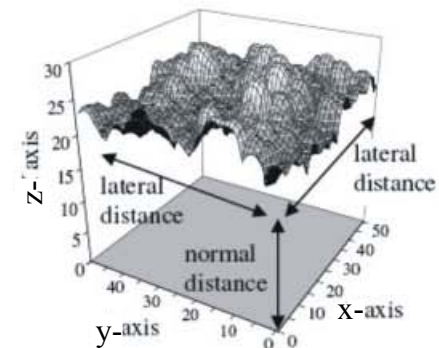
# Surface Roughness



## Surface Atoms



## Surface Characterization



M. Lísal et al., *PCCP* **14**, 5164, 2012

M. Lísal et al., *J. Chem. Phys.* **139**, 014704, 2013

21-22 April 2015



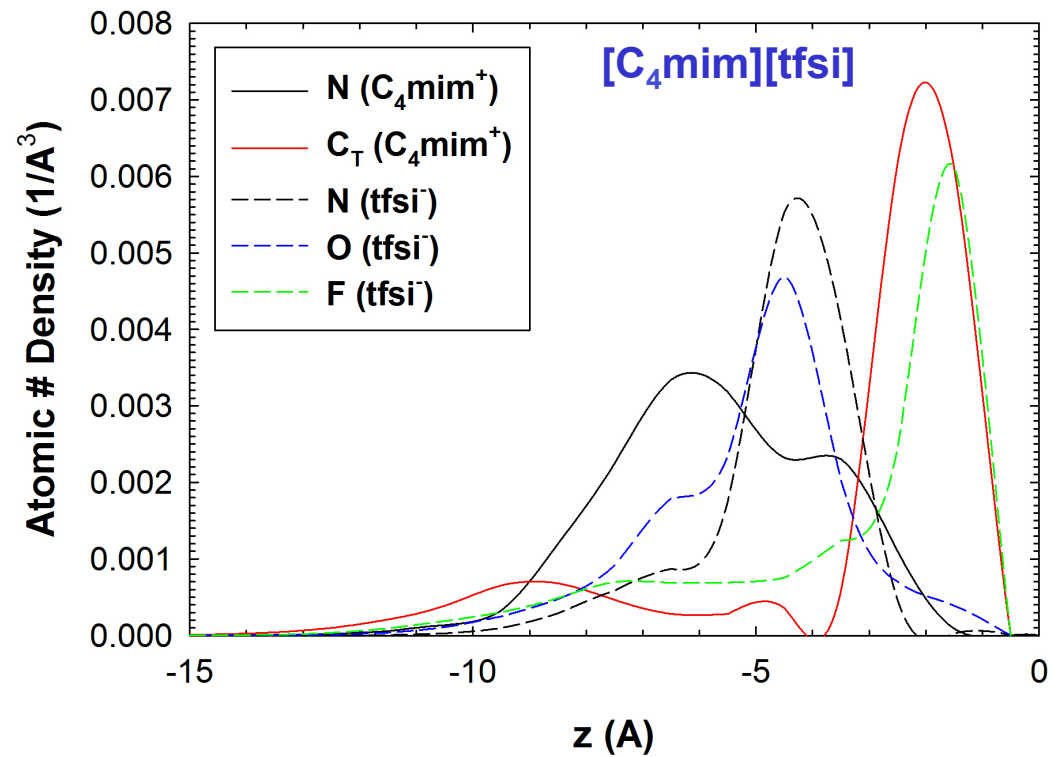
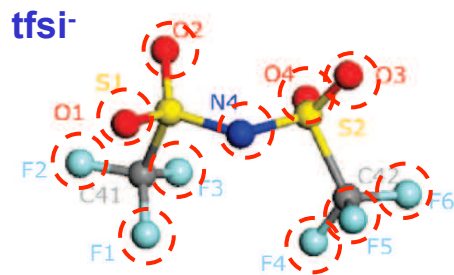
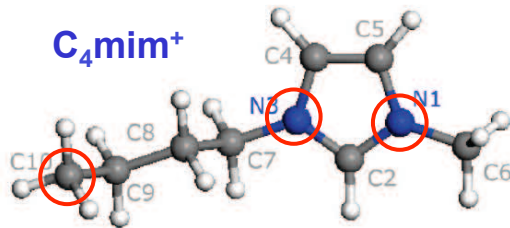
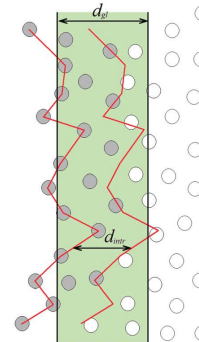
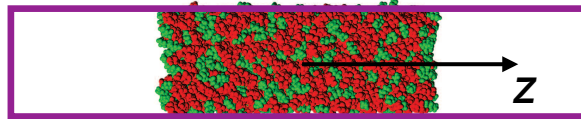
# Summary of Findings #1



- **an enhancement of ions density in the interfacial layer**
- **surplus of the cations wrt the anions in the interfacial layer**
- **the cation surplus is balanced by surplus of the anions in the sub-interfacial layer**
- **the cation surplus increases with increase in the cation tail length**
- **surface of the interfacial layer is smoother than the surface of the sub-interfacial layer which indicates a closer packing of the cations in the interfacial layer**
- **roughness of the surfaces of the interfacial and sub-interfacial layers increased with the increase in the cation tail length which suggests a less dense packing of ions in the surface layers of systems with longer cation tails**



# Interfacial-Layer Atomic Density Profiles

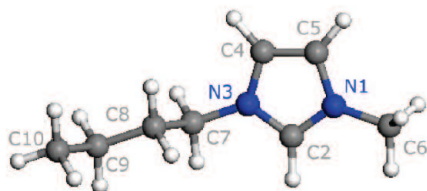




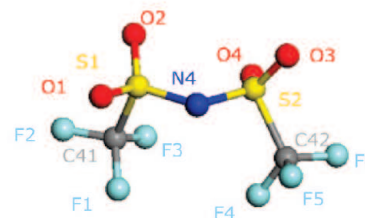
# Preferential Orientations of Ions



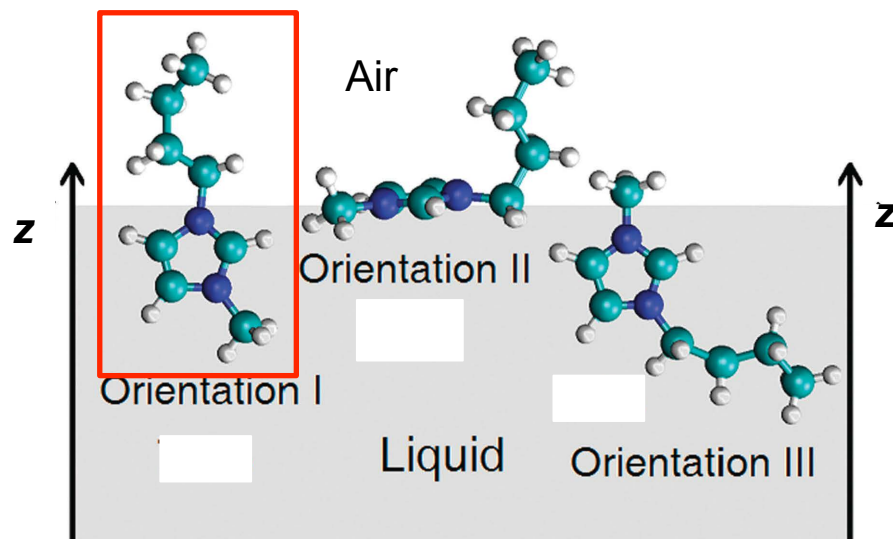
$C_n\text{mim}^+$



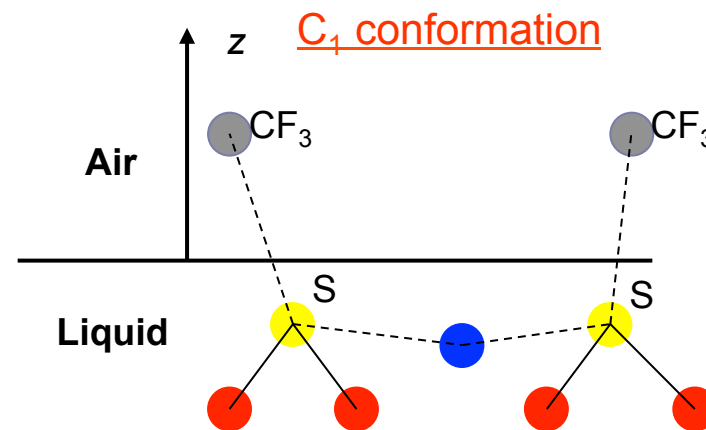
$\text{tfsi}^-$



## Preferential Orientations



## Preferential Orientation



!!! Bulk:  $C_2$  conformation !!!



## Effect of Anion Size

$\gamma_{\text{MeOH}} \sim 22 \text{ mN/m}$

$\gamma_{\text{acetone}} \sim 24 \text{ mN/m}$

$\gamma_{\text{n-alkanes}} \sim (21-23) \text{ mN/m}$

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| RTILs                      | $\gamma^{sim}$<br>(mN/m) | $\gamma^{exp}$<br>(mN/m) | $\Delta\gamma$<br>(%) |
|----------------------------|--------------------------|--------------------------|-----------------------|
| [C <sub>4</sub> mim][tfsi] | 33.6(33)                 | 32.66(70)                | 3.0                   |
| [C <sub>4</sub> mim][Br]   | 39.4(29)                 | 43                       | -8.0                  |

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$\gamma_{\text{water}} \sim 72 \text{ mN/m}$

## Effect of Alkyl Tail Length

| RTILs                      | $\gamma^{sim}$<br>(mN/m) | $\gamma^{exp}$<br>(mN/m) | $\Delta\gamma$<br>(%) |
|----------------------------|--------------------------|--------------------------|-----------------------|
| [C <sub>4</sub> mim][tfsi] | 33.6(33)                 | 32.66(70)                | 3.0                   |
| [C <sub>6</sub> mim][tfsi] | 33.4(37)                 | 31.45(70)                | 6.0                   |
| [C <sub>8</sub> mim][tfsi] | 35.5(40)                 | 31.03(59)                | 12.6                  |

M. Lísal et al., *PCCP* **14**, 5164, 2012

M. Lísal et al., *J. Mol. Liq.* **189**, 85, 2014

M. Lísal et al., *J. Chem. Phys.* **139**, 014704, 2013



## Summary of Findings #2



- cation and anion have preferential orientations simultaneously in the interfacial layer
- cation alkyl tails protrude to the vapor phase irrespective of the tail length
- a less dominant cation orientation with the alkyl tails pointing toward the liquid
- $\text{CF}_3$  groups of  $\text{tfsi}^-$  point toward the vapor phase,  $\text{C}_1$  conformation
- cation and anion show only a weak preferential orientation in the sub-interfacial layer
- surface tension decreases with increasing cation alkyl tails



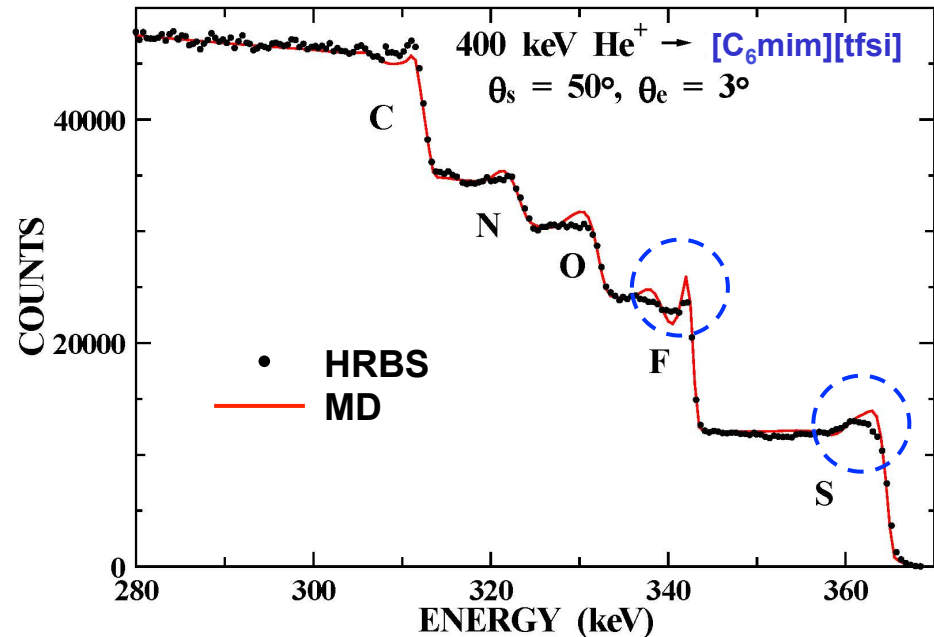
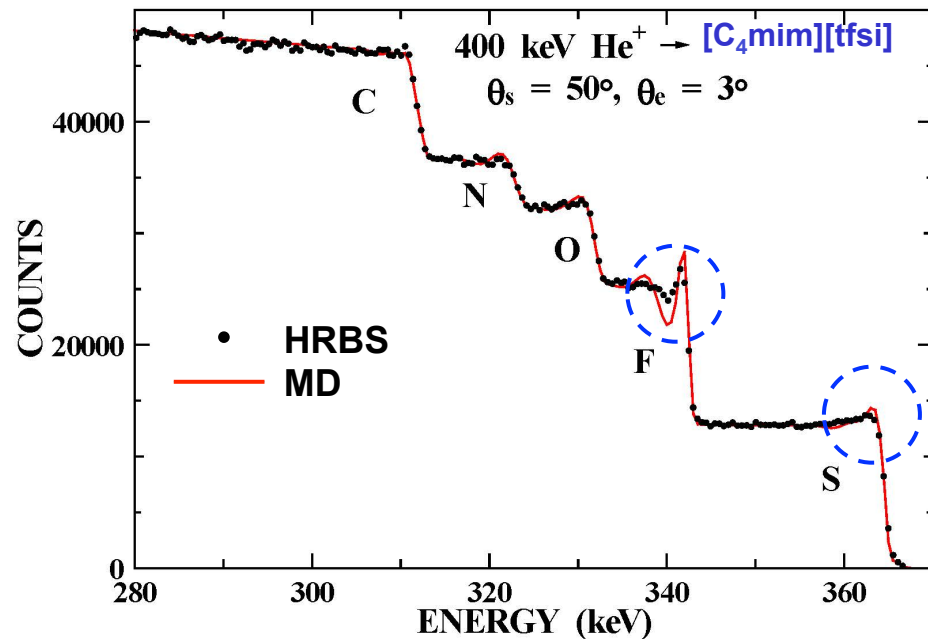


# Comparison with HRBS



- High-Resolution Rutherford Backscattering Spectroscopy (HRBS)<sup>1</sup> (Kenji Kimura's group, Kyoto, Japan)
- HRBS provides elemental depth profiles with a sub-nanometer depth resolution

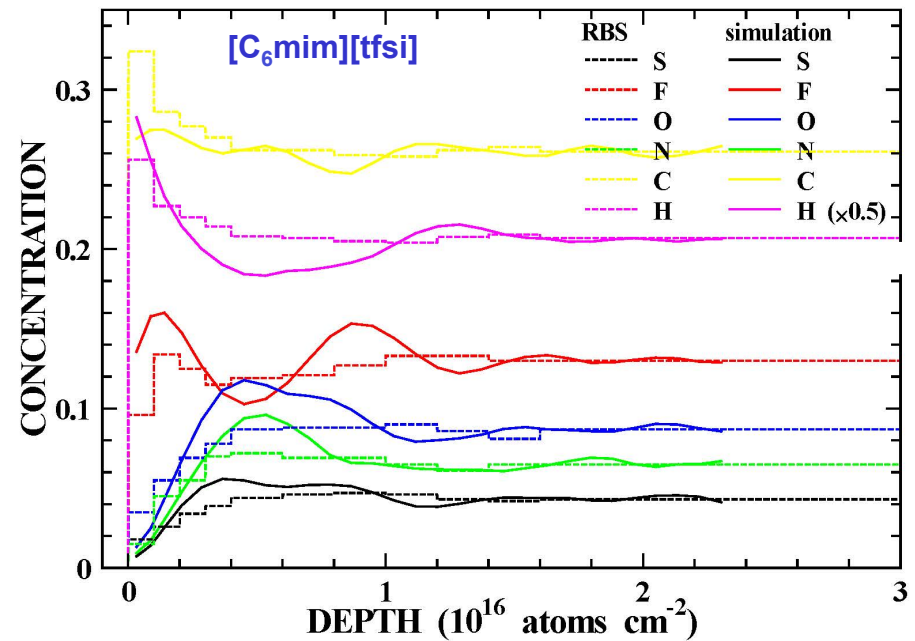
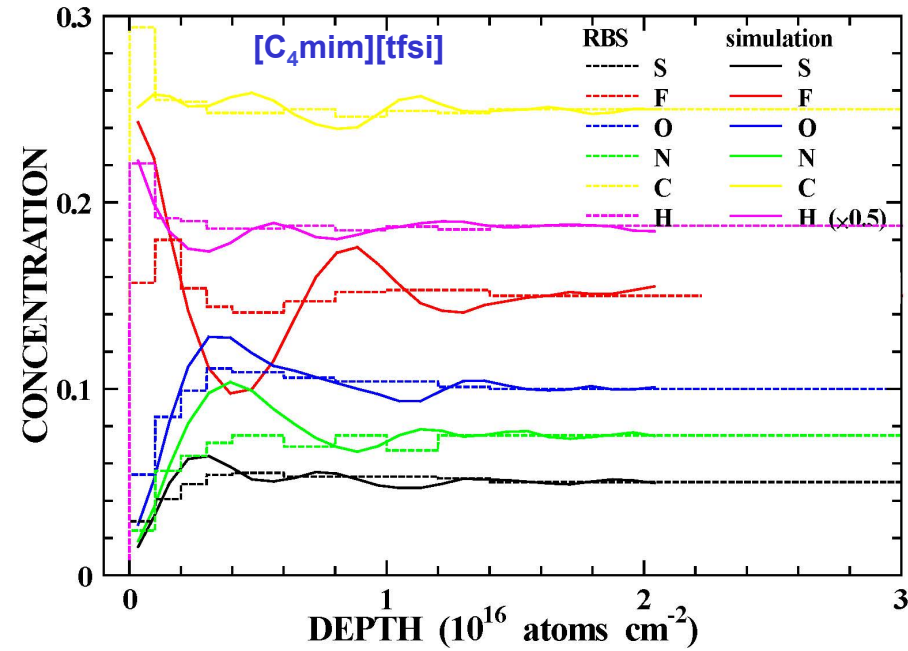
## HRBS Energy Spectra



<sup>1</sup>K. Nakajima et al., *J. Chem. Phys.* **133**, 044702, 2010



## HRBS Elemental Depth Profiles





## Summary: HRBS vs. MD [ $C_n$ mim][tfsi]



### Rutherford Backscattering Spectroscopy

- The surface is shared almost equally by  $C_n$ mim<sup>+</sup> and tfsi<sup>-</sup>.
- At the surface, both  $C_n$ mim<sup>+</sup> and tfsi<sup>-</sup> have preferential orientations simultaneously.
- The alkyl tails of  $C_n$ mim<sup>+</sup> protrude to the air irrespective of the tail length.
- CF<sub>3</sub> groups of tfsi<sup>-</sup> point toward the air.
- tfsi<sup>-</sup> in the sub-interfacial layer show a weak preferential orientation.

### Molecular Dynamics Combined with Intrinsic Analysis

- The surface shows surplus of  $C_n$ mim<sup>+</sup> wrt tfsi<sup>-</sup> and the surplus increases with the cation tail length; the surplus is balanced in the sub-interfacial layer.
- At the surface, both  $C_n$ mim<sup>+</sup> and tfsi<sup>-</sup> have preferential orientations simultaneously.
- The alkyl tails of  $C_n$ mim<sup>+</sup> protrude to the air irrespective of the tail length; a less dominant orientation with the alkyl tails pointing toward the liquid.
- CF<sub>3</sub> groups of tfsi<sup>-</sup> point toward the air.
- Both  $C_n$ mim<sup>+</sup> and tfsi<sup>-</sup> in the sub-interfacial layer show a weak preferential orientation.



# Orientations of Ions

